

Name _____

Chemistry Worksheet
Naming & Formula Writing (Ionic)

Instructions: Write the formulas &/or the names for the compounds listed below

- | | | | |
|--------------------------|-------|--|-------|
| 1. Sodium nitrate | _____ | 26. Aluminum chloride | _____ |
| 2. Calcium carbonate | _____ | 27. Iron (III) hydroxide | _____ |
| 3. Magnesium oxide | _____ | 28. Sodium acetate | _____ |
| 4. Ammonium sulfide | _____ | 29. calcium hydroxide | _____ |
| 5. Lead (II) sulfate | _____ | 30. sodium iodate | _____ |
| 6. Sodium cyanide | _____ | 31. Nickel (II) nitrate | _____ |
| 7. Potassium hydroxide | _____ | 32. Iron (II) chloride | _____ |
| 8. Silver chloride | _____ | 33. Magnesium bromide | _____ |
| 9. Iron (III) hydroxide | _____ | 34. Ammonium nitrate | _____ |
| 10. Potassium hydroxide | _____ | 35. Silver bromide | _____ |
| 11. Tin (IV) perchlorate | _____ | 36. $\text{Al}(\text{OH})_3$ | _____ |
| 12. Potassium carbonate | _____ | 37. NH_4I | _____ |
| 13. Silver nitrate | _____ | 38. Li_2CO_3 | _____ |
| 14. Sodium iodide | _____ | 39. CuSO_4 | _____ |
| 15. Ammonium hydroxide | _____ | 40. KCN | _____ |
| 16. Potassium iodate | _____ | 41. $\text{Pb}(\text{ClO})_2$ | _____ |
| 17. Lead (IV) oxide | _____ | 42. BaS | _____ |
| 18. Ammonium hydroxide | _____ | 43. ZnSO_4 | _____ |
| 19. Barium sulfate | _____ | 44. $\text{Pb}(\text{CH}_2\text{COO})_2$ | _____ |
| 20. barium chloride | _____ | 45. $\text{Ca}(\text{NO}_3)_2$ | _____ |
| 21. Cobalt (II) chloride | _____ | 46. $\text{Fe}_2(\text{CO}_3)_3$ | _____ |
| 22. Sodium carbonate | _____ | 47. NH_4IO_3 | _____ |
| 23. Calcium oxide | _____ | 48. CaCl_2 | _____ |
| 24. Lead (II) nitrate | _____ | 49. NaF | _____ |
| 25. Tin (II) chloride | _____ | 50. $\text{Cu}(\text{NO}_3)_2$ | _____ |

Ion (_____) Formula Chart

Name and Charges of Some Common Ions

Oxidation #: _____

+ ions are called

Mostly

| 1+ | 2+ | 3+ |
|---|---|---|
| ⇒ Ammonium (NH ₄) ¹⁺ ⇒ Cesium Cs ⁺ ⇒ Copper (I) Cu ⁺ ⇒ Gold (I) Au ⁺ ⇒ Hydrogen H ⁺ ⇒ Lithium Li ⁺ ⇒ Potassium K ⁺ ⇒ Rubidium Rb ⁺ ⇒ Silver Ag ⁺ ⇒ Sodium Na ⁺ | ⇒ Barium Ba ²⁺ ⇒ Beryllium Be ²⁺ ⇒ Cadmium Cd ²⁺ ⇒ Calcium Ca ²⁺ ⇒ Chromium (II) Cr ²⁺ ⇒ Cobalt (II) Co ²⁺ ⇒ Copper (II) Cu ²⁺ ⇒ Iron (II) Fe ²⁺ ⇒ Lead (II) Pb ²⁺ ⇒ Magnesium Mg ²⁺ ⇒ Mercury (I) (Hg ₂) ²⁺ ⇒ Mercury (II) Hg ²⁺ ⇒ Nickel Ni ²⁺ ⇒ Strontium Sr ²⁺ ⇒ Tin (II) Sn ²⁺ ⇒ Zinc Zn ²⁺ | ⇒ Aluminum Al ³⁺ ⇒ Chromium (III) Cr ³⁺ ⇒ Cobalt (III) Co ³⁺ ⇒ Gallium Ga ³⁺ ⇒ Gold (III) Au ³⁺ ⇒ Iron (III) Fe ³⁺ |
| | | 4+ |
| | | ⇒ Lead (IV) Pb ⁴⁺ ⇒ Tin (IV) Sn ⁴⁺ |

- ions are called

Mostly

| 1- | 2- | 3- |
|---|---|---|
| ⇒ Acetate (C ₂ H ₃ O ₂) ¹⁻ ⇒ Bromate (BrO ₃) ⁻ ⇒ Bromide Br ⁻ ⇒ Chlorate (ClO ₃) ⁻ ⇒ Chloride Cl ⁻ ⇒ Chlorite (ClO ₂) ⁻ ⇒ Cyanide (CN) ⁻ ⇒ Dihydrogen Phosphate (H ₂ PO ₄) ⁻ ⇒ Fluoride F ⁻ ⇒ Hydrogen Carbonate (HCO ₃) ⁻ AKA—Bicarbonate ⇒ Hydrogen Sulfate (HSO ₄) ⁻ ⇒ Hydroxide (OH) ⁻ ⇒ Iodate (IO ₃) ⁻ ⇒ Iodide I ⁻ ⇒ Nitrate (NO ₃) ⁻ ⇒ Nitrite (NO ₂) ⁻ ⇒ Permanganate (MnO ₄) ⁻ | ⇒ Carbonate (CO ₃) ²⁻ ⇒ Chromate (CrO ₄) ²⁻ ⇒ Dichromate (Cr ₂ O ₇) ²⁻ ⇒ Hydrogen Phosphate (HPO ₄) ²⁻ ⇒ Oxide O ²⁻ ⇒ Oxalate (C ₂ O ₄) ²⁻ ⇒ Peroxide (O ₂) ²⁻ ⇒ Selenide Se ²⁻ ⇒ Sulfate (SO ₄) ²⁻ ⇒ Sulfide S ²⁻ ⇒ Sulfite (SO ₃) ²⁻ ⇒ Tartrate (C ₄ H ₄ O ₆) ²⁻ ⇒ Telluride Te ²⁻ ⇒ Thiosulfate (S ₂ O ₃) ²⁻ | ⇒ Borate (BO ₃) ³⁻ ⇒ Nitride N ³⁻ ⇒ Phosphate (PO ₄) ³⁻ ⇒ Phosphide P ³⁻ |

Polyatomic Ion: _____